

# Calibration of Thermometers for Marine Applications



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Thermometers for marine applications are usually based on thermistors:  
temperature-sensitive **semiconducting** ceramic devices

### **Advantages:**

- 1. very high sensitivity (typically ten times that of PRT)
- 2. small size (some smaller than 0.2 mm)
- 3. fast time constants (some as short as a few milliseconds)
- 4. High resistance values

### **Disadvantages:**

- very high non-linearity,
- a limited temperature range,
- a risk of high self-heating due to the sensing current

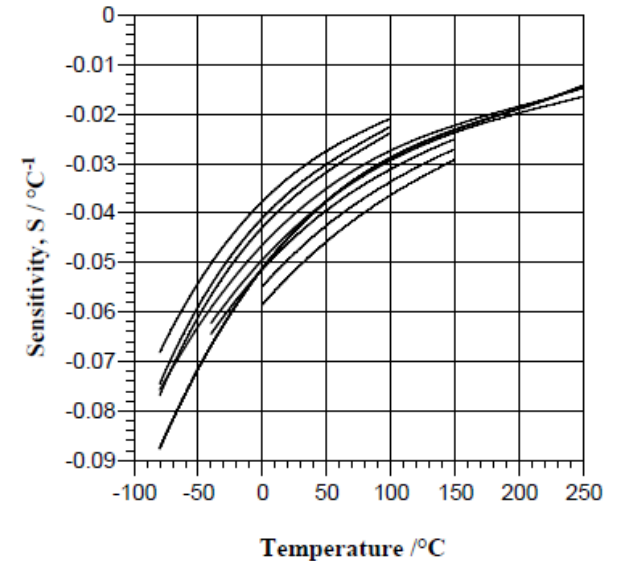
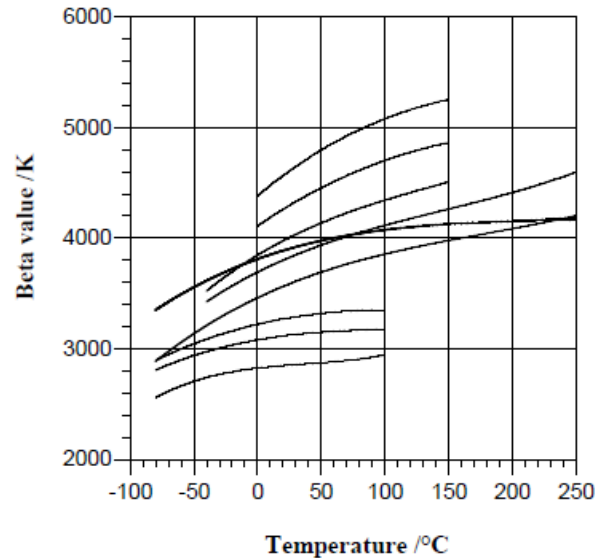
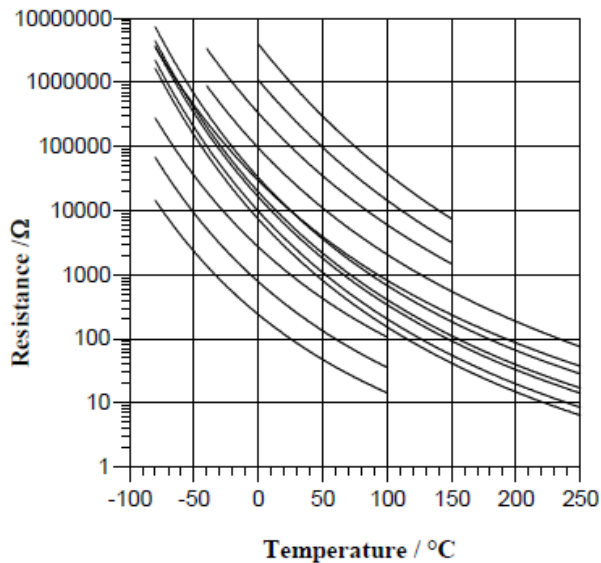


$$R(T) = R(T_0) \exp \left[ \beta \left( \frac{1}{T} - \frac{1}{T_0} \right) \right]$$

$$S = \frac{1}{R} \frac{dR}{dT} = \frac{d \ln(R)}{dT} = -\frac{\beta}{T^2}$$

$T_0$ : is a reference temperature, often 25 °C

$\beta$ : is a characteristic of the thermistor material



We measure  $R \rightarrow T$  ??

By the calibration of the thermistors at different temperatures

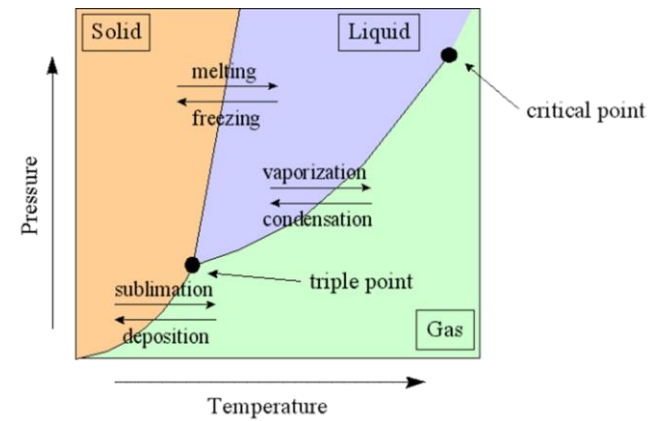


## Calibration of Sea-Bird 35 in fixed points:

H<sub>2</sub>O: 0.01 °C

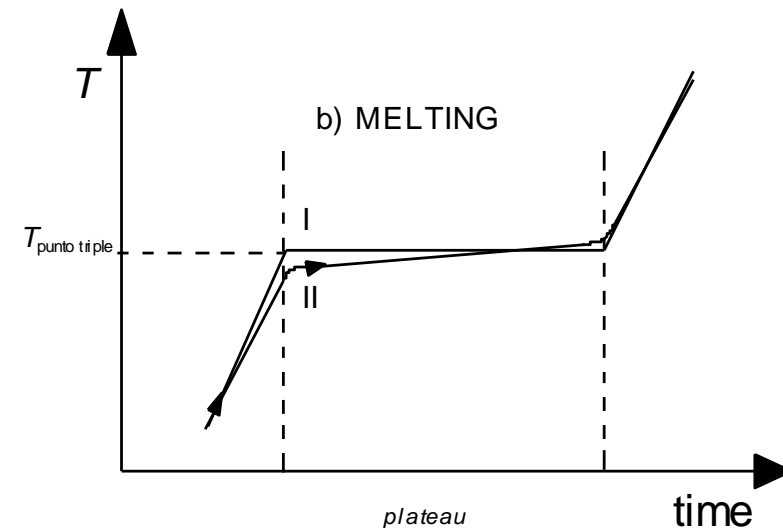
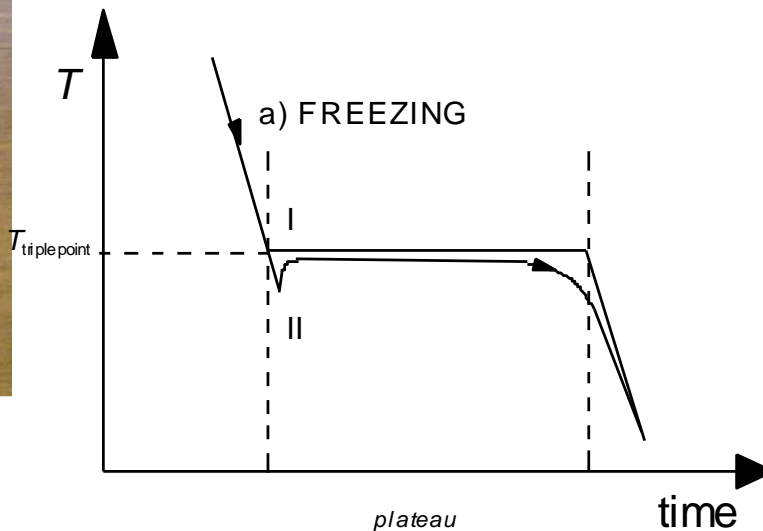
Ga: 29.7646 °C

### Melting and freezing plateaus



The phase transition plateaus are the period of time in which a substance experience a phase transition:

- From liquid to solid: freezing plateau
- From solid to liquid: melting plateau



The objective is to get as much **flat** plateaus and as much **longer** plateaus as possible

# ITS-90. Defining Fixed points

## The Triple Point of Water

A triple point of water cell contains ice, water and water vapour.

In practice the thermometer needs to be immersed in a place at the triple point temperature. This is achieved with an ice mantle surrounding the thermometer well, with a thin water interface between the ice and the well. This allows the water-ice interface to be in close thermal contact with the thermometer.

To prepare the triple point:

1. the thermometric well is filled with a refrigerant (dry ice, cold nitrogen gas, metal rod cooled in liquid nitrogen...).
2. Once the ice mantle is large enough, the refrigerant is removed and the cell is maintained in an appropriate isothermal enclosure (i.e. water bath).
3. The thin water interface between the ice and the thermometric well by inserting a metal rod at ambient temperature in the thermometric well.



# ITS-90. Fixed points

## Metal Fixed Points



In most cases, the metal sample is contained in a high purity **graphite crucible**, being the thermometric well also made in graphite.

A volume of 100 ml to 250 ml is usually required. **An immersion** of the thermometer of at least 20 cm is needed for optimal realizations.

The graphite crucible is assembled in a **glass or quartz tube** (depending on the temperature), and a second thermometric well made of glass or quartz is also inserted in the graphite crucible.

In the case of the Ga cell, the crucibles are made of PTFE and Stainless Steel or glass is used for mercury.

# ITS-90. Defining Fixed points

## Metal Fixed Points. Realization.

The fixed point cell is inserted into a **furnace** to induce the phase transition.

The typical furnace is a tube of around 400 mm to 600 mm long.

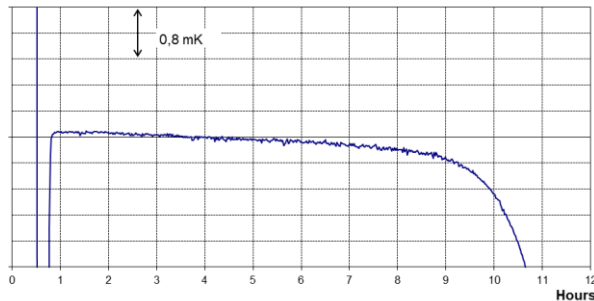
The essential furnace requirements are the **temperature uniformity and stability**.

Furnaces: three zone furnaces and heat pipe furnaces

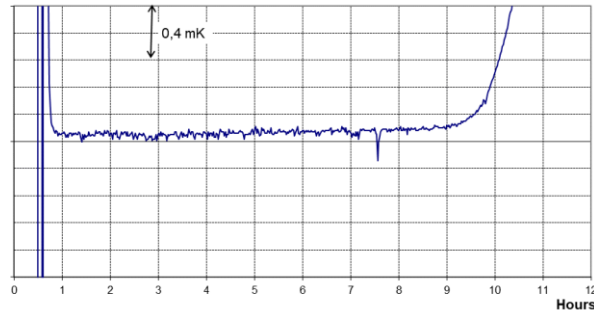
The usual realization procedure of a freezing plateau is, roughly:

- The cell is placed in the furnace with a thermometer in the thermometric well. Then, the temperature of the furnace is raised to about 5 K above the melting point to assure a complete melt.
- Once the melt is completed the furnace is set to about 1 K or 2 K below the freezing point.
- Once the thermometer indicates that a freezing plateau is reached, a cold glass or quartz rod (depending on the temperature) should be inserted in the thermometric well to create a thin solid layer between the melted metal and the thermometric well.

FIXED POINT OF INDIUM (156,598 5 °C)  
FREEZING PLATEAU



FIXED POINT OF GALLIUM (29,764 6 °C)  
MELTING PLATEAU



## Calibration of a thermistor:

A thermistor resistance measurement usually involves a voltage measurement, then the accuracy is limited by the resolution of the voltmeter

Self-heating: 
$$\Delta T_{sh} = I^2 R(T) (\rho_{int} + \rho_{ext}) = \frac{V^2}{R(T)} (\rho_{int} + \rho_{ext}),$$

Depending on the temperature range is better to work at  $V$  constant or  $I$  constant

Reducing the error due to self-heating by the calibration at the same conditions of use

Influence of environmental conditions ( $T$ ): conduction of the leads

	H2O, mK	Ga, mK	
<b>UNCERTAINTIES DUE TO THE REFERENCE SYSTEM</b>			
	0.041 28	0.098 47	
<b>UNCERTAINTY COMPONENTS DUE TO THE SB 35</b>			
Resolution	0.016 79	0.119 67	average of the noise in measurements
Repeatability	0.084 86	0.034 12	maximum of the reading's difference at several fixed points plateaus
Environmental conditions (conduction effect)	0.946 20	0.287 97	Maximum variation of the readings in a fixed point with changed environmental conditions ( $T$ )
Standard uncertainty $u =$	1.0	0.3	mK
Expanded uncertainty, $U(k=2) =$	1.9	0.7	mK



## Calibration of a thermistor:

High sensitivity implies high non-linearity:

Calibration curve:

$$R(T) = R(T_0) \exp \left[ \beta \left( \frac{1}{T} - \frac{1}{T_0} \right) \right],$$

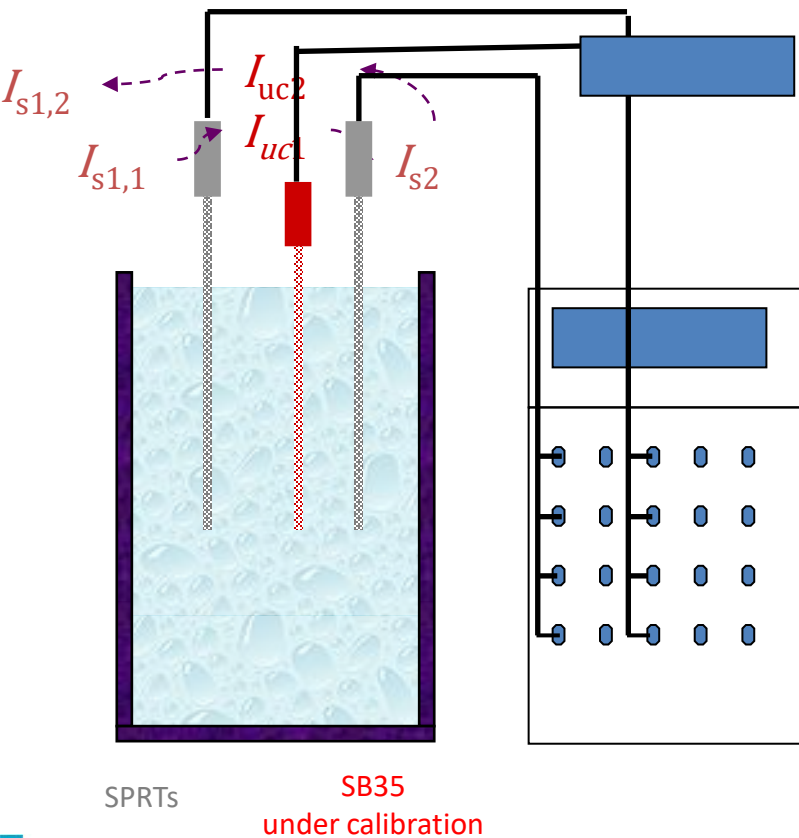
$$\frac{1}{T} = A + B \ln(R/R_0) + C [\ln(R/R_0)]^2 + D [\ln(R/R_0)]^3 + \dots, \quad \ln(R/R_0) = A + \frac{B}{T} + \frac{C}{T^2} + \frac{D}{T^3} + \dots,$$

With only two calibration points, (H<sub>2</sub>O and Ga) the uncertainty of the calibration curve is high

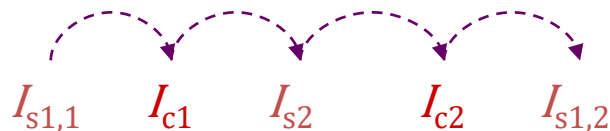
$$u_T^2 = \sum_{i=1}^N \left[ l_i^2(T) \left( \frac{T}{T_i} \right)^{6-2N} \left( u_{T_i}^2 + \frac{T_i^4}{\beta^2} \frac{u_{R_i}^2}{R_i^2} \right) \right] + \frac{T^4}{\beta^2} \frac{u_R^2}{R^2},$$

so more calibration points are needed: **Calibration by comparison in baths**

# Calibration by Comparison. Two reference thermometers.



Reading cycle for each calibration point:



The use of two standards and this reading cycle allows to check the stability and uniformity in temperature of the isothermal enclosure in each calibration point by using the readings the standards:

**Stability check:**

$$I_{s1,1} - I_{s1,2} < \text{stability of isothermal enclosure (previously determined)}$$

**Uniformity check:**

$$(I_{s2} - (I_{s1,1} + I_{s1,2})/2) < \text{uniformity of isothermal enclosure (previously determined)}$$

These checks allow a better assurance of the calibration quality. Uniformity is usually the dominant uncertainty component

	values/mK	probability distribution	Contribution to the combined uncertainty/ mK	
<b>UNCERTAINTIES DUE TO THE REFERENCE SYSTEM</b>				
Reference thermometers calibration	1.5	normal	0.75	
Resistance bridge	1.0	normal	0.50	
<b>UNCERTAINTIES DUE TO THE CALIBRATION BATH</b>				
stability	2.0	normal	1.15	
homogeneity	2.8	rectangular	1.62	
<b>UNCERTAINTY COMPONENTS DUE TO THE SB 35</b>				
Resolution	1.0	rectangular	0.58	average of the noise in measurements
Repeatability	0.2	rectangular	0.10	maximum of the reading's difference at several fixed points plateaus
Environmental conditions (conduction effect)	1.6	rectangular	0.90	Maximum variation of the readings in a fixed point with changed environmental conditions ( $T$ )
Standard uncertainty $u =$			2.43	mK
Expanded uncertainty, $U(k=2) =$			4.9	mK

Reference temperature, °C	Reading SB35, °C	Correction SB35, °C	Expanded Uncertainty ( $k = 2$ ), °C
-0.003	-0.003	0.000	0.005
5.092	5.094	-0.002	0.017
10.091	10.090	0.001	0.005
15.215	15.217	-0.002	0.005
18.150	18.148	0.002	0.005
20.146	20.145	0.001	0.005
22.164	22.164	0.000	0.005
25.124	25.124	0.000	0.005
28.096	28.093	0.002	0.005
30.077	30.076	0.001	0.005
-0.003	-0.003	0.000	0.005



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# Thank you Merci



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